





















All balanced net ionic equations when reduced to the smallest common stoichiometric number is the same for the neutralization of all acids: eg. sulfuric acid, H_2SO_4 (aq), nitric acid HNO₃, phosphoric acid H_3PO_4 and all others.

A) True $X H^{+}_{(aq)} + X OH^{-}_{(aq)} \rightarrow X H_2O_{(l)}$ B) False

Answer All balanced net ionic equations when reduced to the smallest common stoichiometric number is the same for the neutralization of all acids: eg. sulfuric acid, H_2SO_4 (aq), nitric acid H_1NO_3 , phosphoric acid H_3PO_4 and all others. A) True $X H^{+}_{(aq)} + X OH^{-}_{(aq)} \rightarrow X H_2O_{(l)}$ B) False











Answer

<u>1</u>Na₂SO_{4(aq)}+ <u>2</u>Ag(NO₃)_{3(aq)} → <u>1</u>Ag₂SO_{4(s)}+ <u>2</u>NaNO_{3(aq)} SO₄²⁻_(aq)+ <u>2</u>Ag⁺_(aq)→ Ag₂SO_{4(s)} The balanced net ionic equation for the reaction of sodium sulfate and silver nitrate contains which of the following species? A) 2 Na⁺(aq)

B) 2 NO₃⁻(aq) C) 2 Ag⁺(aq) D) 2 AgNO₃(aq)

E) All of the above

QUESTION

Observing a cloudy water insoluble precipitate $Al_2(CO_3)_3(s)$, predict the ions and coefficients for the reacting ions that would produce the precipitate and complete the following net ionic equation:

 $- - + - - \rightarrow Al_2(CO_3)_3 (s)$

A. 2 AlCl₃(aq) + 3 Na₂CO₃(aq) also include 6 NaCl(aq) on right

- B. $3 \text{ Al}^{3+}(aq) + 2 \text{ CO}_{3}^{2-}(aq)$
- C. 2 Al³⁺(aq) + 3 CO₃²⁻(aq)
- D. 2 Al³⁺(aq) + 6 Cl⁻(aq) + 3 CO₃²⁻(aq) + 6 Na⁺(aq)

ANSWER Observing a cloudy water insoluble precipitate $Al_2(CO_3)_3(s)$, predict the ions and coefficients for the reacting ions that would produce the precipitate and complete the following net ionic equation: $----+ --- \rightarrow Al_2(CO_3)_3(s)$ A. 2 AlCl₃(aq) + 3 Na₂CO₃(aq) also include 6 NaCl(aq) on right B. 3 Al³⁺(aq) + 2 CO₃²⁻(aq)

C. 2 Al³⁺(aq) + 3 CO₃²⁻(aq) D. 2 Al³⁺(aq) + 6 Cl⁻(aq) + 3 CO₃²⁻(aq) + 6 Na⁺(aq)

